1. _____ The atomic # of Carbon is 6 which means it has 6 protons and 6 electrons.
2. _____ The # of neutrons = mass # - atomic #.
3. _____ Isotopes are the same elements with different numbers of electrons.
4. _____ The following are correct electric symbols: Q = charge, I = voltage, E = energy.
5. _____ A neutral object is repelled by a charged object.
6. _____ A negative object is repelled by another negative object.
7. _____ The density of an object is calculated by: mass x volume.
8. _____ M = D/V.
9. _____ A Bohr (shell) diagram for nitrogen would have 2 electrons in the first shell and 7 electrons in the second shell.
10. _____ The second shell of an atom can hold a maximum of 8 electrons.
11. _____ Outer electrons of an atom are called "valence" electrons.
12. _____ Using the "cross-down method" the correct formula for sodium oxide would be Na2O3.
13. _____ Using the "cross-down method" the correct formula for barium chloride would be BaCl2.
14. _____ During ionic bond formation, a metal atom loses electrons (which becomes negatively charged) to a non-metal atom (which becomes positively charged).
15. _____ In ionic bond formation between sodium and chlorine, sodium become Na+ and chlorine becomes Cl- and the 2 ions then attract each other.
16. _____ In question #16 above, both Na and Cl achieve full outer shells of electrons.
17. _____ In covalent bonding, ions are formed that share electrons.
18. _____ In drawing electrical circuits, the negative terminal of a battery is the long vertical line and the positive terminal is the short vertical line.
19. _____ Electrons leave the negative terminal and return to the positive terminal of a battery.
20. _____ Examples of electrical "loads" are bulbs and heating coils.
21. _____ Electrical loads have high resistance that hinder the passage of protons through them.
22. _____ When an ebonite rod is rubbed with fur, the rod becomes negatively charged as it gains protons.
23. _____ If an ebonite rod rubbed with fur touches a neutral object, the rod becomes negatively charged.
24. _____ After touching the ebonite rod in #31 above, the ball will be attracted to the ebonite rod.
25. _____ A neutral object is drawn with equal numbers of positive and negative charges.
26. _____ A glass rod has a greater pull on electrons than a plastic bag.
27. _____ A glass rod has a greater pull on electrons than a plastic bag.
28. _____ A heterogeneous mixture has more than one visible particle. An example is water.
29. _____ A homogeneous solution would be apple juice.
30. _____ An example of a pure substance is human blood.
31. _____ The release of light, heat, and sound indicates a physical change.
32. _____ According to particle theory, all particles are at rest until heated.
33. _____ The bottom of a thundercloud is positively charged and induces a negative charge on the ground.
34. _____ When an object is discharged, excess charge (electrons) enter or exit it to "ground".
35. _____ To measure E/Q use a voltmeter, to measure Q/E use an ammeter.
36. _____ The following are correct valences: alkali metals=1, alkaline earth metals=3, halogens=7.
37. _____ There are practice quizzes on zerobio that will help you study for the exam!